

Name _____
Per. _____

Chemistry Test – Molecular Geometry

1. The polarity of a molecule is a direct result of all of the following EXCEPT
 - a. the presence of polar bonds
 - b. the shape of the molecule.
 - c. the types of atoms in the molecule.
 - d. the number of multiple bonds in the molecule.

2. Draw the Lewis structure for the following molecules.

a. Cl_2

b. HCN

c. C_3H_8

3. Which molecule has the greatest bond energy, O_2 , N_2 , Cl_2 ?

4. State the basic assumptions of the VSEPR theory.

5. Explain how a molecule can have polar bonds but be non-polar.